**Review Sheet for Wednesday’s Quiz**

1) A 10.0 L container has a pressure of 2.00 atmospheres and contains 7.95 moles of carbon dioxide. Given this information, what is the temperature of the gas?

2) 7.8 L of a gas is at a pressure of 1.5 atm in a closed bag at a temperature of 21 degrees Celsius. What is the new volume of the bag if the pressure remains the same but the temperature is increased to 28 degrees Celsius.

3) If a truck has an internal temperature of 0.719 atm when the temperature is 10.0 degrees Celsius, what will the pressure be a few hours later if the pressure increases to 0.995 atm?

4) What happens at the triple point of a phase diagram? How is this different from the critical point?

5) How are boiling and vaporization different from one another?

6) Given what you know about phase diagrams, why is it impossible for a gas to exist at absolute zero?